

**A. Rule of REDOX reaction**

What is Redox Reaction?

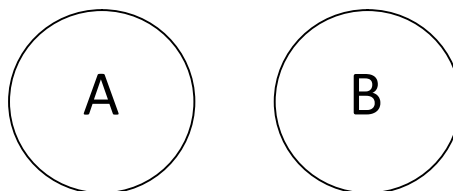
Redox is one kind of \_\_\_\_\_ reaction which is very \_\_\_\_\_ in our daily life.

Example:

**Batteries,** \_\_\_\_\_**Red ox** represents **2 chemical reactions** \_\_\_\_\_ and \_\_\_\_\_**Redox:** \_\_\_\_\_ and \_\_\_\_\_ **MUST** occur at the same time.Remark: **These 2 terms “Oxidation” and “Reduction” are quite misleading****“Oxidation”** may not related to Oxygen, **“Reduction”** is gaining electron during the reaction**Very Basic Principle of Redox!**2 chemical species: A and B: **A** tends to **lose** electron, **B** tends to **gain** electrons**When A meets B,** \_\_\_\_\_ **Reaction occurs**

A: Some Rich Men like to  
\_\_\_\_\_ money.

B: Some Women like to  
\_\_\_\_\_ money.

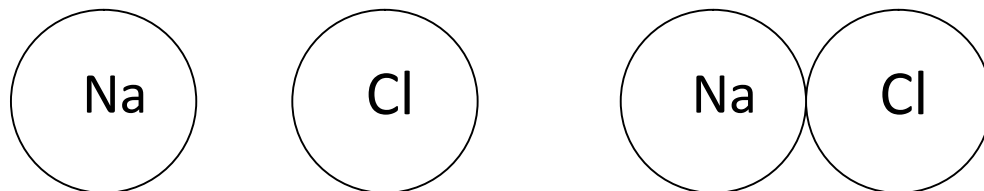


Electron Flows		
Charge of species		
Is oxidized / reduced?		
Reducing agent (RA) Oxidizing agent(OA)		
Oxidation number (Charge of AN atom)		
Type of Reaction		

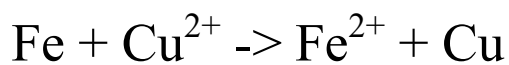
More Example:

2 atoms: Na and Cl : Na tends to \_\_\_\_\_ electron, Cl tends to \_\_\_\_\_ electrons

When Na meets Cl, \_\_\_\_\_ Reaction occurs



Electron Flows		
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Type of Reaction		



Fe

$\text{Cu}^{2+}$

Electron Flows		
Charge of species		
Is oxidized / reduced?		
Reducing agent (RA) Oxidizing agent(OA)		
Oxidation number (Charge of AN atom)		
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Half equations