

Calor sensível

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“

Calor é energia trocada (**ganha ou perdida**)

Temperatura é a medida da vibração das moléculas **do corpo**

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Tipos de calor?



Tipos de calor?



→ Varia temperatura

→ Muda de fase

Tipos de calor?



$$Q = m \cdot c \cdot \Delta T$$

$$Q = m \cdot L$$

$$Q = m \cdot c \cdot \Delta T$$

$$Q = m \cdot c \cdot \Delta T$$

Quantidade
de calor
[cal]

$$Q = m \cdot c \cdot \Delta T$$

Quantidade
de calor
[cal]

massa
[g]

$$Q = m \cdot c \cdot \Delta T$$

Quantidade
de calor
[cal]

massa
[g]

Variação de
temperatura
[°C]

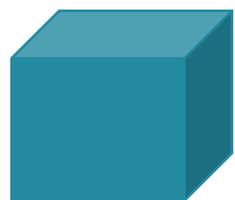
Calor específico (c)

$$Q = m \cdot c \cdot \Delta T$$

[cal/g°C]

Calor específico (c)

$$Q = m \cdot c \cdot \Delta T$$



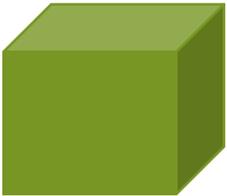
$c = 0,8 \text{ cal/g}^\circ\text{C}$

$[\text{cal/g}^\circ\text{C}]$

Calor específico (c)

$$Q = m \cdot c \cdot \Delta T$$

 $c = 0,8 \text{ cal/g}^\circ\text{C}$

 $c = 0,4 \text{ cal/g}^\circ\text{C}$

[cal/g[°]C]

Calor específico (c)

$$Q = m \cdot c \cdot \Delta T$$

água

$$C_{(s)} = 0,5 \text{ cal/g}^\circ\text{C}$$

$$C_{(L)} = 1,0 \text{ cal/g}^\circ\text{C}$$

$$C_{(v)} = 0,48 \text{ cal/g}^\circ\text{C}$$

[cal/g[°]C]

Calor específico (c)



Calor específico (c)



Capacidade térmica (C)

$$Q = m \cdot c \cdot \Delta T$$

Capacidade térmica (C)

$$Q = m \cdot c \cdot \Delta T$$

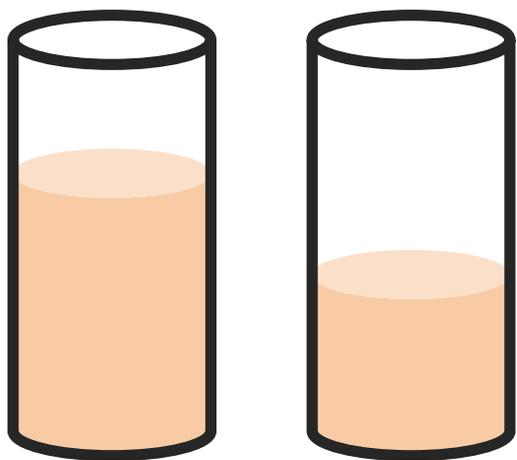
$$Q = C \cdot \Delta T$$

[cal/°C]

Capacidade térmica (C)

$$Q = m \cdot c \cdot \Delta T$$

$$Q = C \cdot \Delta T$$



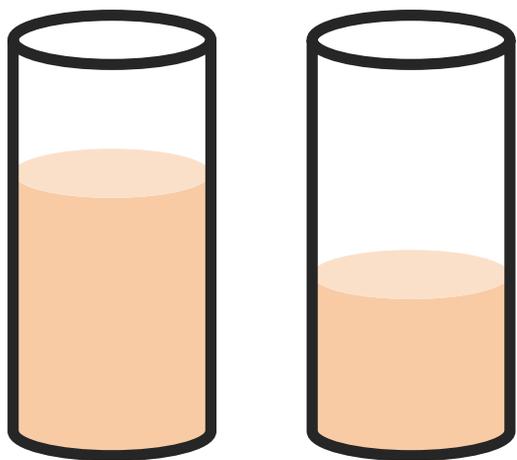
$m > m$

$c = c$

Capacidade térmica (C)

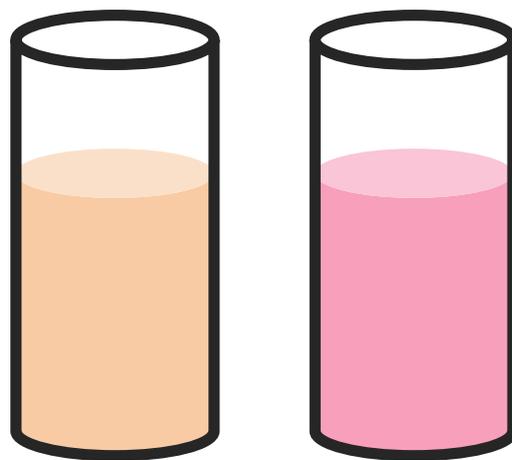
$$Q = m \cdot c \cdot \Delta T$$

$$Q = C \cdot \Delta T$$



$m > m$

$c = c$



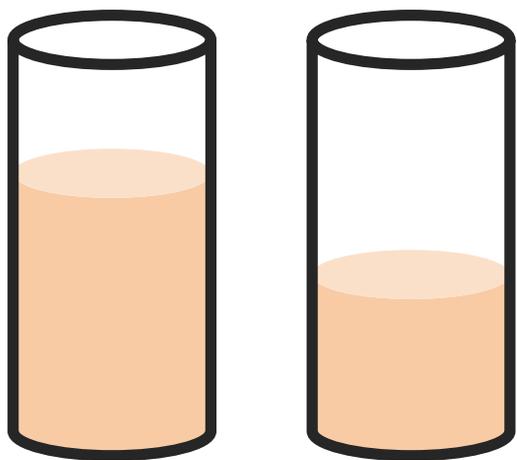
$m = m$

$c < c$

Capacidade térmica (C)

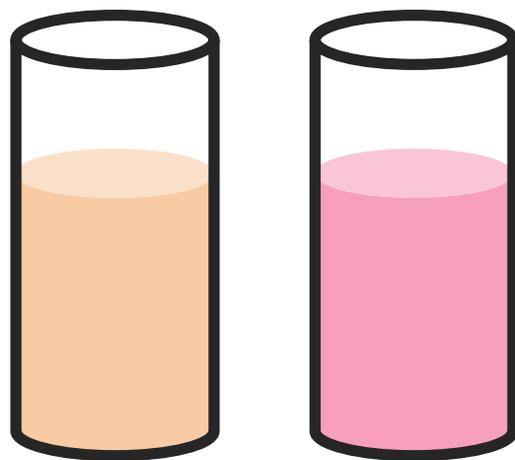
$$Q = m \cdot c \cdot \Delta T$$

$$Q = C \cdot \Delta T$$



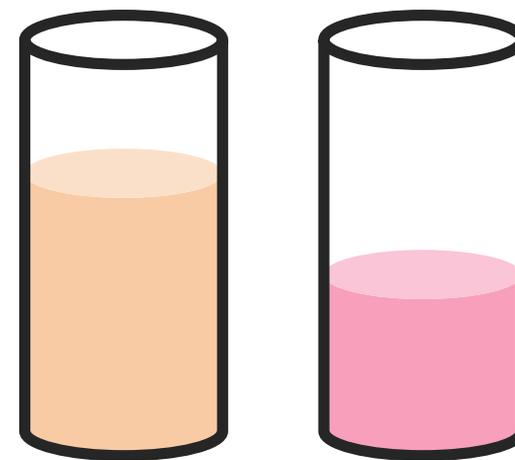
$m > m$

$c = c$



$m = m$

$c < c$



$200g > 100g$

$0,4_{\text{cal/g}^\circ\text{C}} < 0,8_{\text{cal/g}^\circ\text{C}}$

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