

MOCK PAPER 1

CHEMISTRY (Theory)

Class-12th

Time : 3 Hours

Max. Marks : 70

General Instructions:

Read the following instructions carefully and follow them:

- This question paper contains 33 questions. All questions are compulsory.
- This question paper is divided into five sections – Sections A, B, C, D and E.
- Section A – question numbers 1 to 16 are multiple choice type questions. Each question carries 1 mark.
- Section B – question numbers 17 to 21 are very short answer type questions. Each question carries 2 marks.
- Section C – question numbers 22 to 28 are short answer type questions. Each question carries 3 marks.
- Section D – question numbers 29 and 30 are case-based questions. Each question carries 4 marks.
- Section E – question numbers 31 to 33 are long answer type questions. Each question carries 5 marks.
- There is no overall choice given in the question paper. However, an internal choice has been provided in few questions in all the sections except Section – A.
- Kindly note that there is a separate question paper for Visually Impaired candidates.
- Use of calculators is NOT allowed.

You may use the following values of physical constants wherever necessary:

$$c = 3 \times 10^8 \text{ m/s}$$

$$h = 6.63 \times 10^{-34} \text{ Js}$$

$$e = 1.6 \times 10^{-19} \text{ C}$$

$$\mu_0 = 4\pi \times 10^{-7} \text{ T m A}^{-1}$$

$$\epsilon_0 = 8.854 \times 10^{-12} \text{ C}^2 \text{ N}^{-1} \text{ m}^{-2}$$

$$\frac{1}{4\pi\epsilon_0} = 9 \times 10^9 \text{ N m}^2 \text{ C}^{-2}$$

$$\text{Mass of electron } (m_e) = 9.1 \times 10^{-31} \text{ kg.}$$

$$\text{Mass of neutron} = 1.675 \times 10^{-27} \text{ kg.}$$

$$\text{Mass of proton} = 1.673 \times 10^{-27} \text{ kg.}$$

$$\text{Avogadro's number} = 6.023 \times 10^{23} \text{ per gram mole}$$

$$\text{Boltzmann's constant} = 1.38 \times 10^{-23} \text{ JK}^{-1}$$

SECTION – A (16×1 = 16)

Questions No. 1 to 16 are Multiple Choice type questions carrying 1 mark each.

- In case of association, abnormal molar mass of solute will
 - increase
 - decrease
 - remain same
 - first increase and then decrease
- Standard electrode potential for $\text{Sn}^{4+}/\text{Sn}^{2+}$ couple is +0.15 V and that for the Cr^{3+}/Cr couple is –0.74 V. The two couples in their standard states are connected to make a cell. The cell potential will be

- | | |
|-------------|-------------|
| (A) +1.19 V | (B) +0.89 V |
| (C) +0.18 V | (D) +1.83 V |
- The magnetic moment is associated with its spin angular momentum and orbital angular momentum. Spin only magnetic moment value of Cr^{3+} ion (Atomic no. : Cr = 24) is

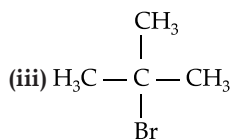
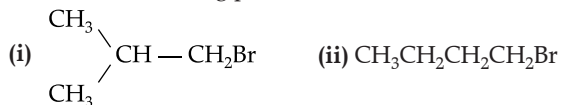
(A) 2.87 B.M.	(B) 3.87 B.M.
(C) 3.47 B.M.	(D) 3.57 B.M.
 - Acidified KMnO_4 oxidises sulphite to

(A) $\text{S}_2\text{O}_3^{2-}$	(B) $\text{S}_2\text{O}_8^{2-}$
(C) $\text{SO}_2(\text{g})$	(D) SO_4^{2-}
 - The correct IUPAC name of $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]^{2+}$ is

(A) Diamminedichloridoplatinum (II)	(B) Diamminedichloridoplatinum (IV)
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- (C) Diamminedichloridoplatinum (O)
(D) Diamminedichloridoplatinate (IV)

6. Arrange the following compounds in increasing order of their boiling points:



The correct order is

- (A) (ii) < (i) < (iii) (B) (i) < (ii) < (iii)
(C) (iii) < (i) < (ii) (D) (iii) < (ii) < (i)
7. Alkyl halides undergoing nucleophilic bimolecular substitution reaction involve
(A) retention of configuration
(B) formation of racemic mixture
(C) inversion of configuration
(D) formation of carbocation
8. Which is the correct order of acid strength from the following?
(A) $\text{C}_6\text{H}_5\text{OH} > \text{H}_2\text{O} > \text{ROH}$
(B) $\text{C}_6\text{H}_5\text{OH} > \text{ROH} > \text{H}_2\text{O}$
(C) $\text{ROH} > \text{C}_6\text{H}_5\text{OH} > \text{H}_2\text{O}$
(D) $\text{H}_2\text{O} > \text{C}_6\text{H}_5\text{OH} > \text{ROH}$
9. The acid formed when propyl magnesium bromide is treated with CO_2 followed by acid hydrolysis is:
(A) $\text{C}_3\text{H}_7\text{COOH}$ (B) $\text{C}_2\text{H}_5\text{COOH}$
(C) CH_3COOH (D) $\text{C}_3\text{H}_7\text{OH}$
10. The best reagent for converting propanamide into propanamine is
(A) excess H_2
(B) Br_2 in aqueous NaOH
(C) iodine in the presence of red phosphorus
(D) LiAlH_4 in ether
11. Which of the following statements is not true about glucose?
(A) It is an aldohexose.
(B) On heating with HI it forms n -hexane.
(C) It exists in furanose form.
(D) It does not give Schiff's test.
12. An unripe mango placed in a concentrated salt solution to prepare pickle, shrivels because
(A) it gains water due to osmosis
(B) it loses water due to reverse osmosis
(C) it gains water due to reverse osmosis
(D) it loses water due to osmosis

For questions number 13 to 16, two statements are given – one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below :

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).

(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).

(C) Assertion (A) is true, but Reason (R) is false.

(D) Assertion (A) is false and Reason (R) is true.

13. **Assertion (A):** $[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_2$ and $[\text{Fe}(\text{H}_2\text{O})_6]\text{Cl}_2$ are example of homoleptic complexes.

Reason (R): All the ligands attached to the metal are the same.

14. **Assertion (A):** The boiling points of alkyl halides decrease in the order :

$\text{RI} > \text{RBr} > \text{RCl} > \text{RF}$

Reason (R): The boiling points of alkyl chlorides, bromides and iodides are considerably higher than that of the hydrocarbon of comparable molecular mass.

15. **Assertion (A):** The boiling points of ethanol is higher than that of methoxymethane.

Reason (R): There is intramolecular hydrogen bonding in ethanol.

16. **Assertion (A):** All naturally occurring α -amino acids except glycine are optically active.

Reason (R): Most naturally occurring amino acids have L-configuration.

SECTION – B

17. (A) Give reasons: [1+1=2]

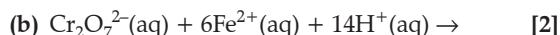
(a) Cooking is faster in pressure cooker than in an open pan.

(b) On mixing liquid X and liquid Y, volume of the resulting solution decreases. What type of deviation from Raoult's law is shown by the resulting solution? What change in temperature would you observe after mixing liquids X and Y?

OR

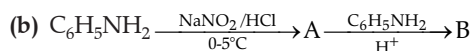
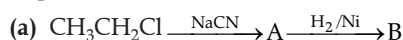
(B) Define Azeotrope. What type of Azeotrope is formed by negative deviation from Raoult's law? Give an example. [2]

18. Complete and balance the following chemical equations:



19. Would you expect benzaldehyde to be more reactive or less reactive in nucleophilic addition reactions than propanal? Justify your answer. [2]

20. Identify A and B in each of the following reaction sequence: [1+1=2]



21. Write the reaction involved when D-glucose is treated with the following reagents : [1+1=2]

- (a) HCN (b) Br_2 water

SECTION – C

22. A solution of glucose (molar mass = 180 g mol^{-1}) in water has a boiling point of 100.20°C . Calculate the freezing point of the same solution. Molal constants for water K_f and K_b are $1.86 \text{ K kg mol}^{-1}$ and $0.512 \text{ K kg mol}^{-1}$ respectively. [3]

23. (a) State the following:

- Kohlrausch law of independent migration of ions and
- Faraday's first law of electrolysis.

- (b) Using E°_{values} of X and Y given below, predict which is better for coating the surface of iron to prevent corrosion and why?

$$\text{Given } E^\circ_{\text{X}^{2+}/\text{X}} = -2.36 \text{ V,}$$

$$E^\circ_{\text{Y}^{2+}/\text{Y}} = -0.14 \text{ V,}$$

$$E^\circ_{\text{Fe}^{2+}/\text{Fe}} = -0.44 \text{ V} \quad [3]$$

24. A certain reaction is 50% complete in 20 minutes at 300 K and the same reaction is 50% complete in 5 minutes at 350 K. Calculate the activation energy if it is a first order reaction.

$$[R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}; \log 4 = 0.602] \quad [3]$$

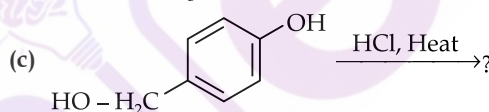
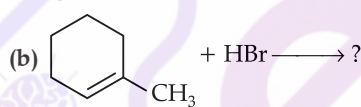
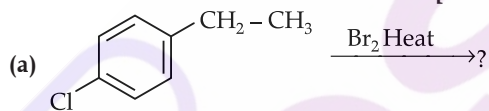
25. The elements of 3d transition series are given as :

Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn [1+1+1=3]

Answer the following :

- Copper has exceptionally positive $E^\circ_{\text{M}^{2+}/\text{M}}$ value, why?
 - Which element is a strong reducing agent in +2 oxidation state and why?
 - Zn^{2+} salts are colourless. Why?
26. (A) Draw the structure of the major monohalo product for each of the following reaction:

[1+1+1=3]



OR

- (B) How do you convert: [3×1=3]

- Chlorobenzene to biphenyl
- Propene to 1-Iodopropane
- 2-bromobutane to but-2-ene.

27. (a) Arrange the following compounds in increasing order of their boiling point: [1+2=3]



- (b) Give plausible explanation for each of the following :

- Aromatic primary amines cannot be prepared by Gabriel Phthalimide synthesis.
- Amides are less basic than amines.

28. (a) What is the difference between native protein and denatured protein? [1+1+1=3]

- (b) Which one of the following is a disaccharide?
Glucose, Lactose, Amylose, Fructose

- (c) Which vitamin is responsible for the coagulation of blood?

SECTION – D

The following questions are case based questions. Read the passage carefully and answer the questions that follow.

29. The rate of a chemical reaction is expressed either in terms of decrease in the concentration of reactants or increase in the concentration of a product per unit time. Rate of the reaction depends upon the nature of reactants, concentration of reactants, temperature, presence of catalyst, surface area of the reactants and presence of light. Rate of reactions is directly related to the concentration of reactant. Rate law states that the rate of reaction depends upon the concentration terms on which the rate of reaction actually depends, as observed experimentally. The sum of powers of the concentration of the reactants in the Rate law expression is called order of reaction while the number of reacting species taking part in an elementary reaction which must collide simultaneously in order to bring about a chemical reaction is called molecularity of the reaction.

Answer the following questions: [1+1=2]

- (i) What is a rate determining step?
(ii) Define complex reaction.
- What is the effect of temperature on the rate constant of a reaction? [1]

OR

- Why is molecularity applicable only for elementary reactions whereas order is applicable for elementary as well as complex reactions? [1]
- The conversion of molecule X to Y follows second order kinetics. If concentration of X is increased 3 times, how will it affect the rate of formation of Y? [1]

30. Phenols undergo electrophilic substitution reactions readily due to the strong activating effect of OH group attached to the benzene ring. Since, the OH group increases the electron density more to *o*- and *p*-positions therefore OH group is *ortho*, *para*-directing. Reimer-Tiemann reaction is one of the examples of aldehyde group being introduced on the aromatic ring of phenol, *ortho* to the hydroxyl group. This is a general method used for the *ortho*-formylation of phenols.

Answer the following questions:

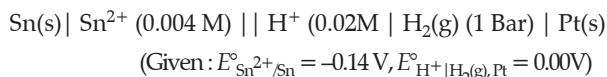
- What happens when phenol reacts with
(i) Br_2/CS_2
(ii) Conc. HNO_3 [2]
- Why phenol does not undergo protonation readily? [1]
- Which is a stronger acid – phenol or cresol? Give reasons. [1]

OR

- Write the IUPAC name of the product formed in the Reimer-Tiemann reaction. [1]

SECTION – E

31. (A) (a) Write the cell reaction and calculate the e.m.f. of the following cell at 298 K : [3+2=5]



- (b) Account for the following :

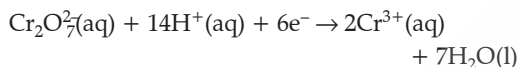
- (i) On the basis of E° values, O_2 gas should be liberated at anode but it is Cl_2 gas which is liberated in the electrolysis of aqueous NaCl .
- (ii) Conductivity of CH_3COOH decreases on dilution.

OR

- (B) (a) Write the anode and cathode reaction and the overall cell reaction occurring in a lead storage battery during its use. [2+3=5]

- (b) Calculate the potential for half-cell containing 0.01 M $\text{K}_2\text{Cr}_2\text{O}_7(\text{aq})$, 0.01 M $\text{Cr}^{3+}(\text{aq})$ and $1.0 \times 10^{-4} \text{ M H}^+(\text{aq})$.

The half cell reaction is



and the standard electrode potential is given as $E^\circ = 1.33 \text{ V}$. [Given : $\log 10 = 1$]

32. (A) Answer the following: [5]

- (a) Low spin tetrahedral complexes are not known.
- (b) Co^{2+} is easily oxidised to Co^{3+} in the presence of a strong ligand [At. No. of Co = 27]
- (c) What type of isomerism is shown by the complex $[\text{Co}(\text{NH}_3)_6][\text{Cr}(\text{CN})_6]$?

- (d) Why a solution of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ is green while a solutions of $[\text{Ni}(\text{CN})_4]^{2-}$ is colourless. (At. No. of Ni = 28)

- (e) Write the IUPAC name of the following complex : $[\text{Co}(\text{NH}_3)_5(\text{CO}_3)] \text{Cl}$

OR

- (B) (a) What is meant by 'Chelate effect' ? Give an example.

- (b) Write the hybridisation and magnetic behaviour of $[\text{Fe}(\text{CN})_6]^{4-}$.

(Atomic number : Fe = 26)

- (c) If $\text{PtCl}_2 \cdot 2\text{NH}_3$ does not react with AgNO_3 , what will be its formula ? [2+2+1=5]

33. (A) (a) Carry out the following conversions :

- (i) Ethanal to But-2-enal
- (ii) Propanoic acid to ethane [2+3=5]

- (b) An alkene A with molecular formula C_5H_{10} on ozonolysis gives a mixture of two compounds B and C. Compound B gives positive Fehling test and also reacts with iodine and NaOH solution. Compound C does not give Fehling solution test but forms iodoform. Identify the compounds A, B and C.

OR

- (B) An organic compound (A) (molecular formula $\text{C}_8\text{H}_{16}\text{O}_2$) was hydrolysed with dilute sulphuric acid to get a carboxylic acid (B) and an alcohol (C). Oxidation of (C) with chromic acid produced (B). (C) on dehydration gives But-1-ene. Identify (A), (B) and (C) and write chemical equations for the reactions involved. [5]